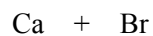
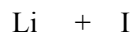


# Chapter 4 Pt 2 Study Guide CP Living

Using electron dot notation, show what happens when the following elements form ionic bonds and make a compound.



What types of elements combine to make an ionic bond? \_\_\_\_\_

Complete the following table:

Element	Symbol	Group Number on Periodic Table	Number of Electrons	Number of Valence Electrons	Number of Core Electrons	Number of Electrons <b>Gained or Lost</b> for Octet	Symbol of Ion
Barium							
Fluorine							
Aluminum							
Sulfur							
Potassium							

## ***Ionic Compounds***

*Use your ion sheet and the criss-cross method to give the formulas of the following ionic compounds.*

Potassium iodide \_\_\_\_\_ Barium chloride \_\_\_\_\_ Lead (II) nitrate \_\_\_\_\_

Copper (II) carbonate \_\_\_\_\_ Magnesium phosphate \_\_\_\_\_ Iron (II) sulfide \_\_\_\_\_

Sodium acetate \_\_\_\_\_ Ammonium hydroxide \_\_\_\_\_ Tin (IV) sulfite \_\_\_\_\_

*Name the following ionic compounds using your ion sheet*

KClO<sub>2</sub> \_\_\_\_\_ Na<sub>2</sub>SO<sub>4</sub> \_\_\_\_\_ Li<sub>2</sub>CO<sub>3</sub> \_\_\_\_\_

CuCl \_\_\_\_\_ Pb(OH)<sub>2</sub> \_\_\_\_\_ Al<sub>2</sub>O<sub>3</sub> \_\_\_\_\_

FeF<sub>2</sub> \_\_\_\_\_ Mg(NO<sub>3</sub>)<sub>2</sub> \_\_\_\_\_ Ni(CN)<sub>3</sub> \_\_\_\_\_